

last name

first name

signature

|                      |                    |                    |                    |                    |                    |                    |                    |                    |                    |                    |                    |                    |                    |                    |                    |                    |                        |
|----------------------|--------------------|--------------------|--------------------|--------------------|--------------------|--------------------|--------------------|--------------------|--------------------|--------------------|--------------------|--------------------|--------------------|--------------------|--------------------|--------------------|------------------------|
| 1<br>1<br>H<br>1.008 |                    |                    |                    |                    |                    |                    |                    |                    |                    |                    |                    |                    |                    |                    |                    |                    | 18<br>2<br>He<br>4.003 |
| 3<br>Li<br>6.941     | 4<br>Be<br>9.012   |                    |                    |                    |                    |                    |                    |                    |                    |                    |                    | 5<br>B<br>10.81    | 6<br>C<br>12.01    | 7<br>N<br>14.01    | 8<br>O<br>16.00    | 9<br>F<br>19.00    | 10<br>Ne<br>20.18      |
| 11<br>Na<br>22.99    | 12<br>Mg<br>24.31  |                    |                    |                    |                    |                    |                    |                    |                    |                    |                    | 13<br>Al<br>26.98  | 14<br>Si<br>28.09  | 15<br>P<br>30.97   | 16<br>S<br>32.07   | 17<br>Cl<br>35.45  | 18<br>Ar<br>39.95      |
| 19<br>K<br>39.10     | 20<br>Ca<br>40.08  | 21<br>Sc<br>44.96  | 22<br>Ti<br>47.87  | 23<br>V<br>50.94   | 24<br>Cr<br>52.00  | 25<br>Mn<br>54.94  | 26<br>Fe<br>55.85  | 27<br>Co<br>58.93  | 28<br>Ni<br>58.69  | 29<br>Cu<br>63.55  | 30<br>Zn<br>65.38  | 31<br>Ga<br>69.72  | 32<br>Ge<br>72.64  | 33<br>As<br>74.92  | 34<br>Se<br>78.96  | 35<br>Br<br>79.90  | 36<br>Kr<br>83.80      |
| 37<br>Rb<br>85.47    | 38<br>Sr<br>87.62  | 39<br>Y<br>88.91   | 40<br>Zr<br>91.22  | 41<br>Nb<br>92.91  | 42<br>Mo<br>95.94  | 43<br>Tc<br>(98)   | 44<br>Ru<br>101.07 | 45<br>Rh<br>102.91 | 46<br>Pd<br>106.42 | 47<br>Ag<br>107.87 | 48<br>Cd<br>112.41 | 49<br>In<br>114.82 | 50<br>Sn<br>118.71 | 51<br>Sb<br>121.76 | 52<br>Te<br>127.60 | 53<br>I<br>126.90  | 54<br>Xe<br>131.29     |
| 55<br>Cs<br>132.91   | 56<br>Ba<br>137.33 | 57<br>La<br>138.91 | 72<br>Hf<br>178.49 | 73<br>Ta<br>180.95 | 74<br>W<br>183.84  | 75<br>Re<br>186.21 | 76<br>Os<br>190.23 | 77<br>Ir<br>192.22 | 78<br>Pt<br>195.08 | 79<br>Au<br>196.97 | 80<br>Hg<br>200.59 | 81<br>Tl<br>204.38 | 82<br>Pb<br>207.20 | 83<br>Bi<br>208.98 | 84<br>Po<br>(209)  | 85<br>At<br>(210)  | 86<br>Rn<br>(222)      |
| 87<br>Fr<br>(223)    | 88<br>Ra<br>(226)  | 89<br>Ac<br>(227)  | 104<br>Rf<br>(267) | 105<br>Db<br>(268) | 106<br>Sg<br>(269) | 107<br>Bh<br>(270) | 108<br>Hs<br>(270) | 109<br>Mt<br>(278) | 110<br>Ds<br>(281) | 111<br>Rg<br>(282) | 112<br>Cn<br>(285) | 113<br>Nh<br>(286) | 114<br>Fl<br>(289) | 115<br>Mc<br>(290) | 116<br>Lv<br>(293) | 117<br>Ts<br>(294) | 118<br>Og<br>(294)     |

|                    |                    |                    |                   |                    |                    |                    |                    |                    |                    |                    |                    |                    |                    |
|--------------------|--------------------|--------------------|-------------------|--------------------|--------------------|--------------------|--------------------|--------------------|--------------------|--------------------|--------------------|--------------------|--------------------|
| 58<br>Ce<br>140.12 | 59<br>Pr<br>140.91 | 60<br>Nd<br>144.24 | 61<br>Pm<br>(145) | 62<br>Sm<br>150.36 | 63<br>Eu<br>151.96 | 64<br>Gd<br>157.25 | 65<br>Tb<br>158.93 | 66<br>Dy<br>162.50 | 67<br>Ho<br>164.93 | 68<br>Er<br>167.26 | 69<br>Tm<br>168.93 | 70<br>Yb<br>173.04 | 71<br>Lu<br>174.97 |
| 90<br>Th<br>232.04 | 91<br>Pa<br>231.04 | 92<br>U<br>238.03  | 93<br>Np<br>(237) | 94<br>Pu<br>(244)  | 95<br>Am<br>(243)  | 96<br>Cm<br>(247)  | 97<br>Bk<br>(247)  | 98<br>Cf<br>(251)  | 99<br>Es<br>(252)  | 100<br>Fm<br>(257) | 101<br>Md<br>(258) | 102<br>No<br>(259) | 103<br>Lr<br>(266) |

constants

$R = 0.08206 \text{ L atm/mol K}$

$R = 8.314 \text{ J/mol K}$

$N_A = 6.022 \times 10^{23} / \text{mol}$

$h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$

$c = 3.00 \times 10^8 \text{ m/s}$

$g = 9.81 \text{ m/s}^2$

conversions

$1 \text{ atm} = 760 \text{ torr}$

$1 \text{ atm} = 101325 \text{ Pa}$

$1 \text{ atm} = 1.01325 \text{ bar}$

$1 \text{ bar} = 10^5 \text{ Pa}$

$^{\circ}\text{F} = ^{\circ}\text{C}(1.8) + 32$

$\text{K} = ^{\circ}\text{C} + 273.15$

conversions

$1 \text{ in} = 2.54 \text{ cm}$

$1 \text{ ft} = 12 \text{ in}$

$1 \text{ yd} = 3 \text{ ft}$

$1 \text{ mi} = 5280 \text{ ft}$

$1 \text{ lb} = 453.6 \text{ g}$

$1 \text{ ton} = 2000 \text{ lbs}$

$1 \text{ tonne} = 1000 \text{ kg}$

$1 \text{ gal} = 3.785 \text{ L}$

$1 \text{ gal} = 231 \text{ in}^3$

$1 \text{ gal} = 128 \text{ fl oz}$

$1 \text{ fl oz} = 29.57 \text{ mL}$

water data

$C_{s,\text{ice}} = 2.09 \text{ J/g } ^{\circ}\text{C}$

$C_{s,\text{water}} = 4.184 \text{ J/g } ^{\circ}\text{C}$

$C_{s,\text{steam}} = 2.03 \text{ J/g } ^{\circ}\text{C}$

$\rho_{\text{water}} = 1.00 \text{ g/mL}$

$\rho_{\text{ice}} = 0.9167 \text{ g/mL}$

$\rho_{\text{seawater}} = 1.024 \text{ g/mL}$

$\Delta H_{\text{fus}} = 334 \text{ J/g}$

$\Delta H_{\text{vap}} = 2260 \text{ J/g}$

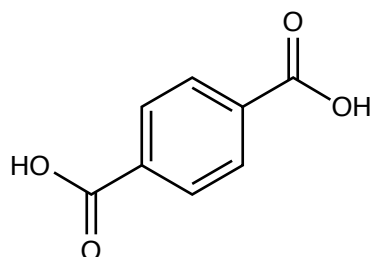
$K_w = 1.0 \times 10^{-14}$

This exam should have exactly 25 questions. Each question is equally weighted at 4 points each. Bubble in your answer choices on the online bubbleseet provided. Your score is based on what you bubble on the bubblesheet and not what is circled on the exam.

---

1. The following structure is a precursor (reactant) to make one of the polymers we have studied. Which polymer (plastic) is it?

- a. PS
- b. nylon
- c. cellulose
- d. proteins
- e. PET
- f. bakelite

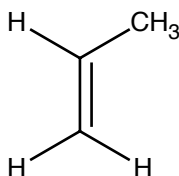


**Explanation:** The structure is terephthalic acid which is mixed with glycol to make PET (polyethylene terephthalate).

---

2. Identify the plastic from the structure of its monomer.

- a. PVC
- b. PP
- c. LDPE
- d. HDPE
- e. PET
- f. PS



**Explanation:** This is propene, or propylene which makes PP (polypropylene).

---

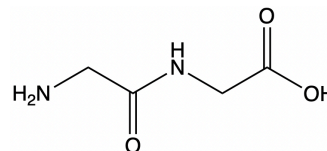
3. What is one major difference between LDPE and HDPE?

- a. HDPE has a lower melting point than LDPE because LDPE is a stronger substance.
- b. Unlike LDPE, HDPE uses a special catalyst to make a completely straight chain with little branching.
- c. Unlike HDPE, LDPE is not made via radical initiated reaction.
- d. LDPE contains less branching which makes it more rigid and a much more crystalline-like substance.

**Explanation:** none

---

4. The image below is a dipeptide, an organic compound derived from two amino acids.



Which of the following statements is true?

- a. The two amino acids that formed this dipeptide reacted in an addition reaction mechanism.
- b. The compound contains an amide link.
- c. The compound contains a primary amine and a tertiary amine.
- d. The compound contains an ester link.

**Explanation:** none

---

5. Five of the Big 6 plastics are all either composed or direct substitutions of which monomer?

- a. ethylene
- b. polychloroethane
- c. diamine
- d. ethyne
- e. methylamine

**Explanation:** Five of the Big 6 plastics are composed of ethene (ethylene), H<sub>2</sub>C=CH<sub>2</sub>.

---

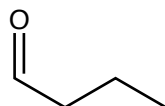
6. Which of the following polymers is the “new” version of the steel-belted radial tire?

- a. silk
- b. bakelite
- c. nylon
- d. polypropylene
- e. kevlar
- f. teflon

**Explanation:** Kevlar is used to make belted radial tires (and bulletproof vests).

7. What is the functional group on propanal (shown below)?

- a. alcohol
- b. ketone
- c. aldehyde
- d. alkene
- e. hydroxyl



**Explanation:** A terminal carbonyl group is an aldehyde.

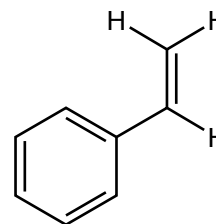
8. Which step of the addition mechanism will create a free radical on a growing polymer chain?

- a. initiation
- b. elimination
- c. propagation
- d. termination
- e. substitution

**Explanation:** The propagation step will grow the polymer chain by one monomer unit and create a radical on the opposite end.

9. Identify the plastic from the structure of its monomer.

- a. PVC
- b. PP
- c. PS
- d. HDPE
- e. PET
- f. LDPE



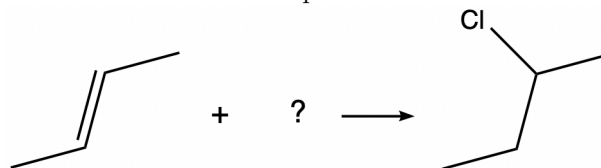
**Explanation:** This is PS.

10. Which feature differentiates RNA nucleotides and allows them to code for the production of different proteins?

- a. the nitrogenous base
- b. the phosphate group
- c. the ribose sugar
- d. the deoxyribose sugar
- e. the R-group

**Explanation:** The nitrogenous base of an RNA nucleotide differentiates it. All RNA nucleotides consist of a ribose sugar and a phosphate group.

11. Given the reaction below, what reactant should be used in order to create the product?



- a. BrCl
- b. CH<sub>3</sub>Cl
- c. HCl
- d. HOCl
- e. Cl<sub>2</sub>

**Explanation:** HCl is added across the double bond

12. Dr. McCord said he still has a lot of old records (LPs) that play on a turntable. What are these records made from?

- a. HDPE
- b. PVC
- c. LDPE
- d. PP
- e. PS
- f. PET

**Explanation:** PVC is used to make vinyl records. The "V" in PVC stands for vinyl. Now very trendy and your hipster friends talk about their lit vinyl collection.

13. Which of the following functional groups contain a carbonyl group?

- I. ketone
  - II. aldehyde
  - III. alcohol
  - IV. carboxylic acid
- a. II and IV
  - b. I and IV
  - c. I, II, and IV
  - d. I, III, and IV
  - e. I, II, III, and IV

**Explanation:** A carbonyl group is a carbon with a double bond to oxygen. Ketones, carboxylic acids, and aldehydes all contain carbonyl groups.

14. Polyethylene can be fashioned into strong milk containers and flimsy plastic bags. Which particular type of polyethylene is used to make the opaque milk containers and why?

- a. LDPE because it is more branched
- b. HDPE because it is a linear polymer
- c. LDPE because it is a linear polymer
- d. HDPE because it is more branched

**Explanation:** HDPE can make liquid containers because it is more linear, creating a more rigid plastic.

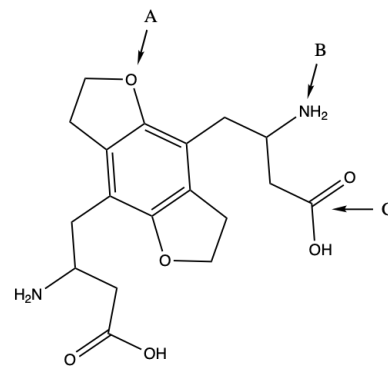
15. Which of the following plastics is NOT formed by an addition mechanism?

- a. PS
- b. HDPE
- c. PVC
- d. LDPE
- e. PET
- f. PP

**Explanation:** PET is formed via condensation.

16. (Part 1 of 2) If you look at it long enough, the following molecule starts to look like an adorable little critter. What is the functional group labeled A?

- a. ketone
- b. aldehyde
- c. amine
- d. carboxylic acid
- e. amide
- f. alcohol
- g. ether



**Explanation:** The R-O-R group is an ether.

17. (Part 2 of 2) What is the functional group labeled B?

- a. amine
- b. alcohol
- c. amide
- d. carboxylic acid
- e. ketone
- f. aldehyde

**Explanation:** The -NH<sub>2</sub> group is an amine.

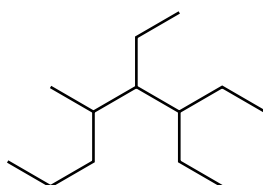
18. (Part 3 of 3) What is the functional group labeled C?

- a. ketone
- b. aldehyde
- c. amide
- d. amine
- e. alcohol
- f. carboxylic acid

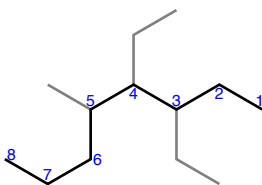
**Explanation:** This is a carboxylic acid.

19. What is the name of this compound?

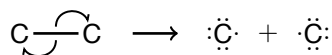
- a. 3,4-diethyl-5-methyloctane
- b. 2-propyl-3,4-diethyloctane
- c. 5,6-diethyl-4-methyloctane
- d. 3,4,5-triethyloctane
- e. 3,4-diethyl-2-propylhexane
- f. 2-propyl-3,4-diethylhexane



**Explanation:** There are 8 carbons in the chain. 3 branch points with two ethyl groups at positions 3 and 4, and a methyl group at position 5. Go alphabetical on ordering, so ethyls are first to get 3,4-diethyl-5-methyloctane.



20. Chlorine is exposed to radiation and breaks into two identical chlorine radicals (shown below). This is an example of...

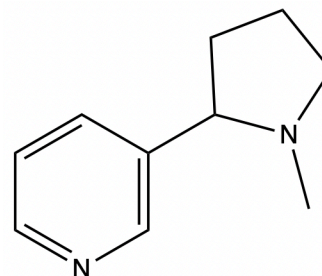


- a. heterolytic cleavage
- b. homoisomeric cleavage
- c. homonuclear fusion
- d. condensation
- e. homolytic cleavage

**Explanation:** The cleavage of a molecule into two identical parts is known as homolytic cleavage.

21. The structure show below is that of a psychoactive stimulant drug that is found in tobacco products. What is the *complete* chemical formula of this drug?

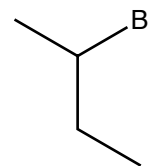
- a.  $\text{C}_{10}\text{H}_{12}\text{N}_2$
- b.  $\text{C}_9\text{H}_{10}\text{N}_2$
- c.  $\text{C}_{10}\text{H}_{10}\text{N}_2$
- d.  $\text{C}_9\text{H}_8\text{N}_2$
- e.  $\text{C}_{10}\text{H}_7\text{N}_2$
- f.  $\text{C}_9\text{H}_{12}\text{N}_2$
- g.  $\text{C}_{10}\text{H}_{14}\text{N}_2$



**Explanation:** There are 10 carbons, 14 hydrogens, and two nitrogens.

22. Name the compound shown below.

- a. 3-bromobutane
- b. 2-bromopentane
- c. 2-bromobutane
- d. 4-bromobutane
- e. 1-bromo-1-methylpropane



**Explanation:** longest chain is 4-carbons and the there is a bromine on position 2. 2-bromobutane.

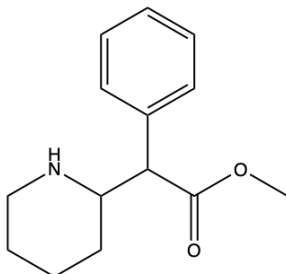
23. Which of the following is a set of two sugar-based polymers?

- a. cellulose and PET
- b. cellulose and starch
- c. wool and rubber
- d. starch and PET
- e. wool and silk

**Explanation:** Cellulose and starch are repeating glucose monomers.

24. Methylphenidate, better known as Ritalin, is a medication that can treat ADHD and narcolepsy. What are two functional groups shown on the ritalin molecule below?

- a. nitrile, ketone
- b. nitrile, ether
- c. primary amine, ester
- d. secondary amine, ether
- e. secondary amine, ester



**Explanation:** The main functional groups on this molecule include a secondary amine and an ester.

---

25. What is the name of the product for the following addition reaction?



- a. 2,3 - dibromohexane
- b. 2 - bromohexane
- c. 1,2 - bromohexane
- d. 4,5 - dibromohexane
- e. 4 - bromohexane

**Explanation:** In the addition reaction, one bromine atom will be placed on each of the carbons sharing a double bond to create the following product: The IUPAC name for this organic compound is 2,3 - dibromohexane because the carbons with the bromine substituents are at carbons 2 and 3, and there are a total of six carbons for the molecule to be a hexane.

---

After you are finished and have all your answers circled, go to the front of the room and then use the QR code show below to pull up the virtual answer page for your exam. Enter the appropriate info plus all your answers - click the SUBMIT button. Double check your choices on the next page. Once you are sure, click the submit button on that page to enter your answers. Make sure you get the confirmation screen (different background color!) and show it to the TA or proctor. After that, turn in your exam and scratch paper. You're free to leave after that.



<https://mccord.cm.utexas.edu/zinc>

\* \* \* Remember your Version Number \* \* \*