□ HI	
□ HBr	
H₂SO₄HAtHCIO₄	
HCIO₄HNO₃NaOH	
□ HCIO □ HCIO ₃	
Question 2	1 pt
What is the pH of a 0.044 M HI solution	
, , Sai a	
Question 3	1 pt
What is the $[OH^{-}]$ when 0.0023 moles of complete dissociation of $Ca(OH)_{2}$.	of Ca(OH) ₂ are placed in 654 mL water? Assume
○ 0.0035 M ○ .0070 M	
2.15 M 3.5 x 10 ⁻⁶ M 12.0 M	
Question 4	1 pt
Use the data here (https://gchem.cm. constants.php) to rank the following wea	utexas.edu/data/section2.php?target=ka-kb- ak acids from weakest to strongest.
HIO CH₃COOH	
HCN HF	
HNO ₂	
HIO < HCN < $CH_3COOH < HNO_2 < HF$ HCN < HIO < $CH_3COOH < HNO_2 < HF$ HNO ₂ < HF < HIO < HCN < CH_3COOH	
○ HF < HNO ₂ < CH ₃ COOH < HCN < HIO	
Question 5	1 pt
A 0.5 M sample of a weak acid, HA ₁ , had acid, HA ₂ , has a pH = 5.66. Which weahled HA ₂ HA ₂	as a pH = 4.24. A 0.5 M sample of another weak k acid has the larger K _a value?
○ HA₁○ Both will have the same value of K_a	
Question 6	1 pt
The generic weak acid HA has a perce concentration. What is the pH?	nt ionization equal to 10.8% at a 0.025 M
Note: Report your answer to two sig fig	s (pH = X.XX)
Question 7	
	1 pt
strong base?	
 Acid + Base → Weak Base + Water Acid + Base → Weak Acid + Water Acid + Base → Acid + Water 	
Acid + Water → Base + SaltBase + Water → Acid + Salt	
Question 8	1 pt
and sodium hydroxide (NaOH) from HV	
	aq) → NaCl(aq) + $H_2O(\ell)$ needed to neutralize 32.0 mL of 0.0291 M HCl?
○ 36.3 mL ○ 20.8 mL	
○ 24.8 mL ○ 33.7 mL ○ 27.1 mL	
○ 49.3 mL	
Question 9	1 pt
(NaOH). The HCl has a concentration of	neutralize a strong acid (HCI) using a strong base of 0.01 M and a volume of 100 mL. The NaOH alsolume of NaOH is needed to fully neutralize the
○ 50 mL	
○ 100 mL ○ 250 mL	
○ 20 mL	
Question 10	1 pt
Barium hydroxide is a strong base that $BaOH_2(aq) \rightarrow Ba^{2+}(aq) + 2OH^{-}(aq)$	dissociates based on the following reaction:
What volume of 0.005 M HCl (strong ad Ba(OH) ₂ solution?	cid) is needed to fully neutralize a 500 mL 0.005 N
○ 1.00 L ○ 500 mL	
○ 1.00 mL	
○ 750 mL ○ 250 mL	
750 mL 250 mL 2.50 L	
750 mL 250 mL 2.50 L Question 11 What is the pH at the equivalence point	1 pt
750 mL 250 mL 2.50 L Question 11 What is the pH at the equivalence point strong base analyte? pH = 7	
750 mL 250 mL 2.50 L Question 11 What is the pH at the equivalence point strong base analyte?	
750 mL 250 mL 2.50 L Question 11 What is the pH at the equivalence point strong base analyte? pH = 7 pH < 7 pH > 7 Question 12	t of a titration involving a strong acid titrant and
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4.65 x 10⁸ moles

1.00 x 10⁻⁷ moles