Comment		1 pts
Common C		
Companies Comp		
	Question 2	1 pts
Process Proc	Hg). How many significant figures are in her recorded number?	(mm of
	infinite	
Company Comp	○ 3	
	terms of matter classification? (check all that apply)	gar in
State	□ liquid	
Caretion of the process of the process of the content of the conte	compound	
The transmission and office in a feet 2 are in a feet an arrangement of the many and office of the second of the s		
Committee Comm	Question 4	1 pts
	strapped together to make a palette. A golf retailer orders 5 palettes of golf balls.	
Administration of regal and policy and the control of the Control		
Comment	Question 5	1 pts
	cylinder is nearby and has 25.6 mL of water in it. The chunk of metal is put into the graduated cylinder and the water line (meniscus) is displaced up to 37.9 mL. Wh	ne
Comment Comm	density of this metal? (answer in g/mi)	
	Question 6	1 pts
Description	Looking carefully at a sidewalk you realize that it is best described as	
State	a heterogeneous mixture	
State		A
Care	What is the atomic mass (aka atomic weight) of potassium?	ı pts
1 25 25 25 25 25 25 25	<u>40.08</u>	
Classification 1 pass of protection 2 pass of prote	39.1032.07	
The septical country on mitters 10, 2, 11, 15 in the allowed is on the part of country and	30.97	
	The sequential counting numbers (1, 2, 3,) for the elements on the periodic tab	
Guestion 9 1 pts Missiones en al source de la composition de la c	known as which of the following? isotopic abundance	
Cuestion 9 All patients on the contraction of the protect of the	o atomic masses	
Extracted and have wellable contractional based on the immunities of the different substances and accordance from the contraction the analysis in the part of the substance and the contraction of the cont		A
Counts on the control of the control	Mixtures can have variable compositions based on the amounts of the different	
Cuestion 10 space of the content of the case of the content or space of the content of space of space of the content of space of the content of space	concentration terms. We chemists have one (and only one) concentration term the have all agreed to use as a standard.	
When all the following statements is make regarding the size of the side in experimental and circumstance and make the control of the size		
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Cuestion 12 1 pts Proce many motes are in 1 46 klograms of sulfur (\$!?) 4.45 more 4.65 more 4.6	What is the molar mass of NH₄Cl?	1 pts
How many motes are in 146 hillograms of sulfur (\$)? 4452 m/vas 455 m/vas 455 m/vas 455 m/vas 450 mores 451 mores 454 more 455 mores 455 more 457 more	○ 50.50 g/mol	1 pts
4.9.22 moles	50.50 g/mol53.49 g/mol49.46 g/mol	1 pts
Set 5 mores Set 50 mores are in 142.5 g methanol, CHyGHT? 4.46 more 4.47 more 5.47 mor	 50.50 g/mol 53.49 g/mol 49.46 g/mol 17.11 g/mol 	
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Sc Mg Na LI Question 16 1 pts Balance the following reaction: —	\$ 50.50 g/mol \$ 53.49 g/mol \$ 17.11 g/mol Question 12 How many moles are in 1.46 kilograms of sulfur (\$)? 46.72 moles . 0.455 moles . 0.910 moles Question 13 How many moles are in 142.5 g methanol, CH ₃ OH? 4.45 mol 4.56 mol No answer text provided. Question 14 How many moles are in 1.85 L H ₂ O? The density of water = 1 kg/ L 3.3.3 mol 103 mol 112.8 mol 0.102 mol	1 pts 1 pts
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 Reactant B will run out while there is still excess A remaining Reactant B will always be the limiting reagent no matter how much of each reactant you begin with There is at least twice the amount of reactant B than A in the beginning of the experiment Reactant A and B will run out simultaneously Question 19 1 pts Consider the following reaction: N₂(g) + 3 H₂(g) → 2 NH₃(g) If 4 moles of N₂ react with 6 moles of H₂, how many moles of NH₃ are formed? 4 moles 8 moles 	\$3.50 g gmal \$4.60 gmal \$4.6	1 pts 1 pts 1 pts 1 pts 1 pts 1 pts
with There is at least twice the amount of reactant B than A in the beginning of the experiment Reactant A and B will run out simultaneously Question 19 1 pts Consider the following reaction: $N_2(g) + 3 H_2(g) \rightarrow 2 NH_3(g)$ If 4 moles of N_2 react with 6 moles of H_2 , how many moles of N_3 are formed? 4 moles 8 moles	© 46.56 gines © 46.66 gines © 46.75 moles © 45.75 moles © 45.7	1 pts 1 pts 1 pts 1 pts 1 pts 1 pts
Question 19 $ 1 \text{ pts} $ Consider the following reaction: $ N_2(g) \ + \ 3 \ H_2(g) \ \rightarrow \ 2 \ NH_3(g) $ If 4 moles of N_2 react with 6 moles of H_2 , how many moles of N_3 are formed?	© 50.50 g/mid © 50.50 g/mid © 40.46 g/mid © 17.11 g/mo Cuestion 12 How many moles are in 1.46 kilograms of surfur (S)? © 40.72 moes © 45.55 moles © 50.50 g/mid A 55 moles © 50.50 g/mid A 55 moles © 50.50 g/mid A 455 mole © 455 moles © 455 moles © 50.50 g/mid A 455 mole © 455 moles © 455 moles © 10.50 moles are in 142.5 g methanol, CH ₂ OH? © 445 mol © 455 mol © 455 moles © 10.50 moles are in 1.85 L H ₂ O? The density of water = 1 kg/L © 10.50 mol © 10.50	1 pts 1 pts 1 pts 1 pts 1 pts 1 pts 1 pts 1 pts
Consider the following reaction: $N_2(g) \ + \ 3 \ H_2(g) \ \rightarrow \ 2 \ NH_3(g)$ If 4 moles of N_2 react with 6 moles of H_2 , how many moles of N_3 are formed? $ \ \ $	Set 96 grave \$5.54 grave \$4.64 grave \$1.711 ghred 1.711 ghred 1.	1 pts
If 4 moles of N ₂ react with 6 moles of H ₂ , how many moles of NH ₃ are formed? O 4 moles O 8 moles	So Se gried So Se gried So Se gried 40 46 gried 71 if girco Question 12 From many moles are in 1.46 kilograms of sulfur (S)? 60 72 m/se 60 72 m/se 60 72 m/se 60 73 m/se 60 74 m/se 610	1 pts 1 pts 1 pts 1 pts 1 pts 1 pts 1 pts
O 8 moles	Sp 50 g privo 65 50 g privo 65 50 g privo 77 11 g pried 20	1 pts 1 pts 1 pts 1 pts 1 pts 1 pts 1 pts
	State grow 12 48 46 you 177-11 gred Question 12 Ches many mose are in 1.46 Morganis of sulfur (8)? 4.87 rows 3.98 Cinuries 3.98 Cinuries 3.91 Cinuries 3.91 Cinuries 4.45 in ord 4.55 in ordes of an unknown metal (M) weights 109.35 g. What is the identity of the information of the identity of t	1 pts 1 pts 1 pts 1 pts 1 pts 1 pts 1 pts
	SA4 seymo State prime State	1 pts 1 pts 1 pts 1 pts 1 pts 1 pts 1 pts 1 pts

Consider the following balanced reaction:

of carbon dioxide formed?

660 g/mol

352 g/mol

704 g/mol

440 g/mol

 $C_2 H_4(g) \ + \ 3O_2(g) \ \to \ 2CO_2(g) \ + \ 2H_2O(\ell)$

When 15 moles of ${\rm O_2}$ are reacted to completion with 8 moles of ${\rm C_2H_4}$, what is the ${\it mass}$