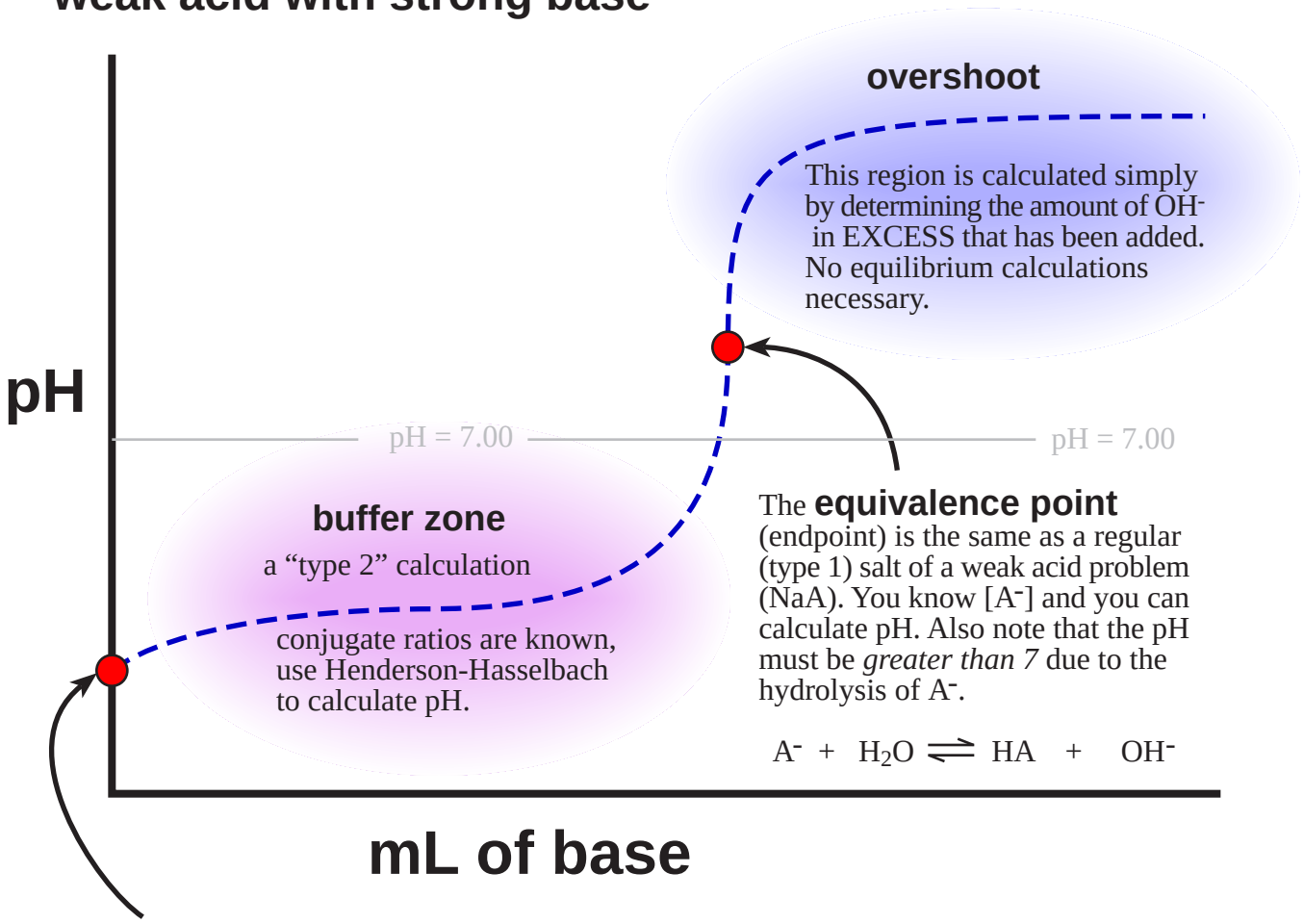
Titration Curve

weak acid with strong base

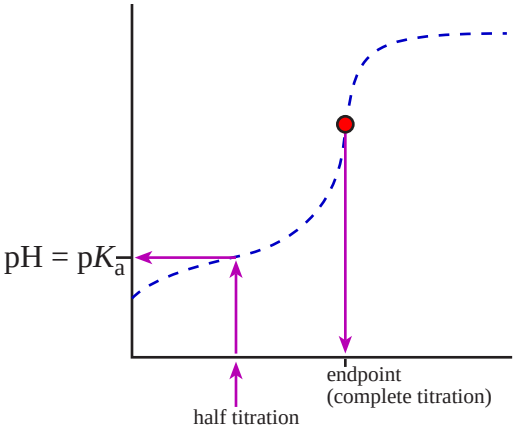


The **START** of the titration is the same as a regular (type 1) weak acid problem. You know K_a and [HA] so you can calculate pH.

$$HA \rightleftharpoons H^+ + A^-$$

The half-way point is important!

After you have determined the equivalence point (endpoint) of the titration, go to half that value. The pH at the half-titration point is equal to the pK_a of the weak acid.



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