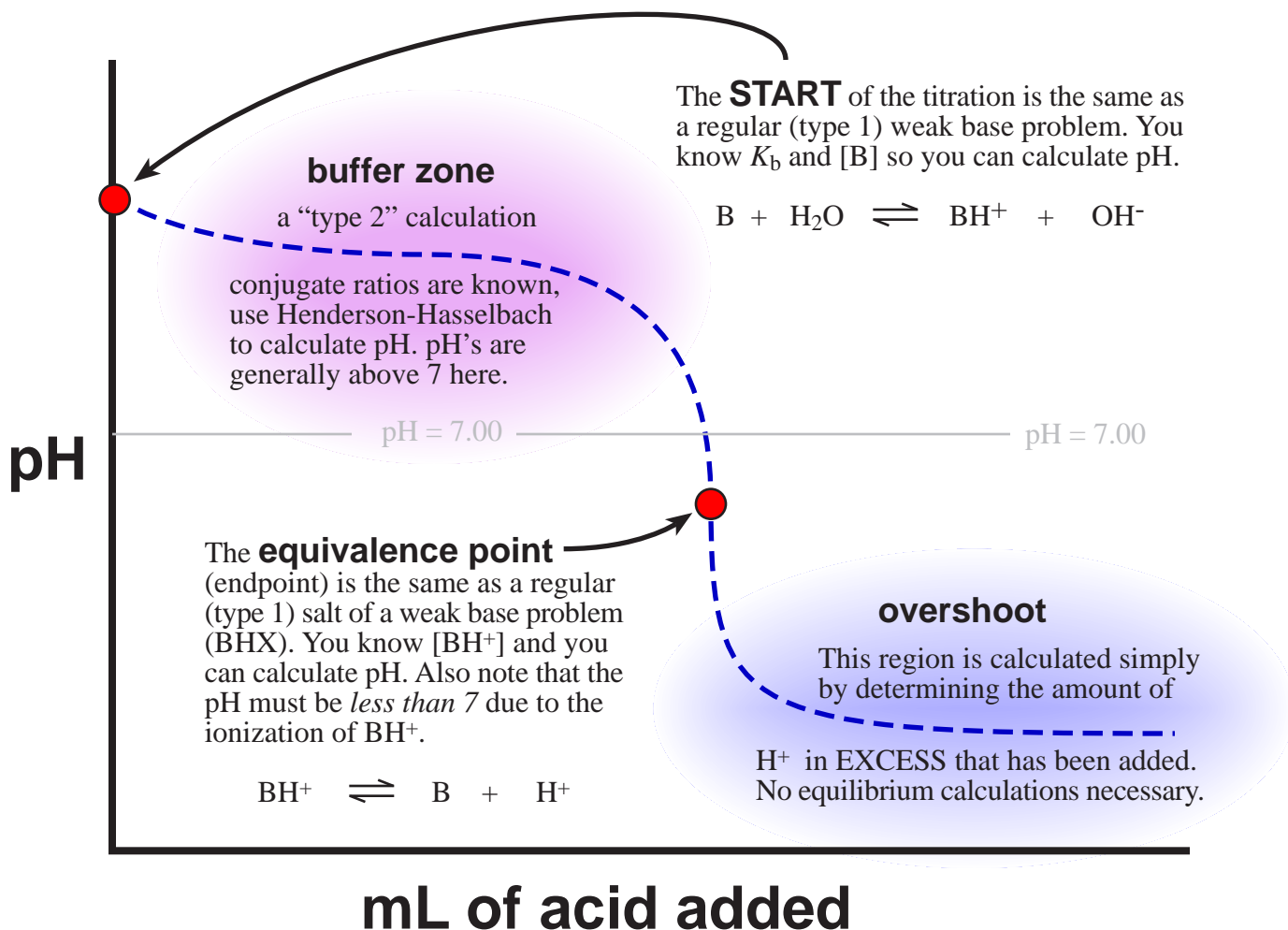


Titration Curve

weak base with strong acid



The half-way point is important!

After you have determined the equivalence point (endpoint) of the titration, go to half that value. The pH at the half-titration point is equal to the pK_a of the weak acid, BH^+ . To get the pK_b of the base (B) you **MUST** subtract the pK_a from 14. The reason for this is that the pOH is actually what equals the pK_b .

$$pK_b = 14 - pK_a$$

